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(b) Divide Molar Mass By Empirical Formula Mass Molar Mass 98.96 G Empirical Formula Mass 49.48 G = 2 = (n) (c) Multiply Empirical Formula By N Obtained Above To Get The Molecular Formula Empirical Formula = CH 2 Cl, N = 2. Hence Molecular Formula Is C 2 H 4 Cl 2. Problem 1.3 Calculate The A 4th, 2024

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. Back Of Chapter Questions . 12.1. Choose The Correct Alternative From The Clues Given At The End Of Each Statement: (a) The Size Of The Atom In Thomson's Model Is _____ The Atomic Size In Rutherford's Model. (much Greater Than/no Different From/much Less Than.) 3th, 2024

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• Standard And Recommended Symbols For Units Are Written In Lower Case Roman (upright) Type, Starting With Small Letters. The Shorter Designations For Units Such As Kg, M, S, Cd, Etc., Are Symbols And Not The Abbreviations. The Unit Names Are Never Capitalised. However, 4th, 2024

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