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(mole-mass, Mass-mole Problems) Stoichiometry Worksheet #2 (mole-mass, Mass-mole Problems) 1. $N_2 + 2O_2 \rightarrow N_2O_4$ A. If 15.0g Of N_2O_4 Was Produced, How Many Moles Of O_2 Were Required? $15.0g N_2O_4 \cdot \frac{1 \text{ Mol } N_2O_4}{92.0g N_2O_4} = 0.163 \text{ Mol } N_2O_4$ B. If 4.0×10^{-3} Moles Of Oxygen Reacted, How Many Grams Of N_2 Were Needed? $4.0 \times 10^{-3} \text{ Mol } O_2 \cdot \frac{1 \text{ Mol } N_2}{2 \text{ Mol } O_2} = 2.0 \times 10^{-3} \text{ Mol } N_2$... Jul 1th, 2024.

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