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Mole Fraction Molality Molarity - University Of Texas At ...

Mole Fraction Sugar Of Solution χ Sugar = $(0.292 \text{ Mol})/(0.292 + 55.6) = .00522 \text{ Mole Fraction Water Of Solution } \chi$ Water = $(55.6 \text{ Mol})/(0.292 + 55.6) = 0.995 \text{ (or } 1-\chi \text{ Sugar)}$ Molality M = (.292 Mol)/(1 Kg) = 0.292 Mol Kg-1 Molarity M = (.292 Mol)/(1 3th, 2024)

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As 1 Atm = 1.013×105 Pa, The Osmotic Pressure Is $\Pi = (0.0036 \times 1.013 \times 105)$ Pa = 365 Pa Hence $C = \Pi / RT = (365 Pa) / ((8.314 M3 Pa K-1 Mol-1) \times (293 K)) = 0.15 Mol M-3 As 1 M3 = 1000 L, This Corresponds To <math>C = 1.5 \times 10$ 1th, 2024

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