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Actual yield = 1.45g Theoretical yield = 5.95g  $\text{KClO}_4 \times 1 \text{ mol}$   
 $122.55 \text{ g} \times 2 \text{ mol} = 245.10 \text{ g}$   $\text{KClO}_4 \times 32.00 \text{ g} \times 1 \text{ mol} = 32.00 \text{ g}$   
 $\% \text{ Yield} = \frac{1.45 \text{ g}}{3.11 \text{ g}} \times 100\% = 46.6\%$   
 $3 \text{ FeCO}_3 + \text{O}_2 \rightarrow 2 \text{ Fe}_2\text{O}_3 + 4 \text{ CO}_2$   
 Percent purity = 62.8%  
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In A Camping Stove? 2. How Many Moles Of  $\text{O}_2$  Molecules Should Be Supplied To Burn 1 Mol Of  $\text{CH}_4$  ...  
 Was At One Time Considered For Use As A Rocket Fuel.  
 ... How Many Formula Units Of Calcium Oxide ( $\text{CaO}$ )  
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 Answers Calcium  $\text{C}_2\text{O}_2 + 3 \text{ Cr}(3)$  When The Potassium  
 Chlorate ( $\text{KClO}_3$ ) Is Heated, It Decomposes To Form  
 Gases  $\text{KCl}$  And Oxygen ( $\text{O}_2$ ).  
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