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Equation With Mostly Whole Numbers: $2 \text{NH}_3 + \frac{7}{2} \text{O}_2 \rightarrow 2 \text{NO}_2 + 3 \text{H}_2\text{O}$. The
Question Asks About The Amount Of Oxygen Reacting With ONE Mole Of Ammonia,
Thus Cut The $\frac{7}{2}$ (3.5) Of Oxygen In Half To 1.75 2. C Balance: CH 3th,
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